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When an ion that is common to the equilibrium system is introduced to the system, the molar solubility of the slightly soluble salt will decrease. Problem no longer requires a stoichiometric relationship between the ionic species. Calculation of S will be based on concentration of the non -common ion.

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to Q_{crit} As confidence level increases, Found in table of Q_{crit} values (Table 7 -5) Value depends on number of measurements AND the confidence level If $Q_{exp} > Q_{crit}$ then reject the outlier. Treat data as if the point never existed Do not use for mean, standard deviation, or any other calculation for the data set!

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In 1993, he received the ACS Division of Analytical Chemistry Award for Excellence in Teaching. He was a Professor of Chemistry at Stanford University. F. James Holler is a Professor of Chemistry and recipient of the Alumni Association Great Teacher Award at the University of Kentucky. He received his Ph.D. from Michigan State University.

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